

Ex 1. How many chloride anions are in your sample?

$$x \text{ Cl}^- = 1.5 \text{ NaCl} = \frac{80.1 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol NaCl}} \times \frac{1 \text{ mol Cl}^-}{1 \text{ mol NaCl}} = 6.02 \times 10^{23} \text{ Cl}^-$$

$$= \frac{80.1 \times 6.02}{58.5} \times 10^{23} \text{ Cl}^-$$

Ex 2. How many hydrogen atoms in your sample?

$$x \text{ H atoms} = 1.5 \text{ sugar} = \frac{90.1 \text{ g sugar}}{342 \text{ g sugar}} \times \frac{1 \text{ mol sugar}}{1 \text{ mol sugar}} \times \frac{22 \text{ mol H atoms}}{1 \text{ mol sugar}} \times \frac{6.02 \times 10^{23} \text{ H atoms}}{1 \text{ mol H atoms}}$$