

Wednesday (4/11/18)

Ms Michener

Chemistry

Academic Chem
Period 2

- * How many formula units are in 0.736 moles of NaCl?
- * What is the percent composition of gallium (III) molybdate?

* define the following:

- > Avogadro's number
- > empirical formula
- > percent(age) composition

*

Finished quiz? calculate number of Na⁺ and Cl⁻ ions in your lab sample. Hint: 1 mole Na⁺ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺
 1 mole Cl⁻ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺

(Lab) record data (measurements) for sugar

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Chemistry

Academic Chem
Period 4

- * What is the mass of 2.44×10^{24} molecules of silicon (IV) hydride?
- * What is the empirical formula of a compound containing the following?

- > 0.890 g K
- > 1.18 g Cr
- > 1.27 g O

* define the following:

- > hydrate
- > mole
- > molar mass

*

Finished quiz? calculate number of Na⁺ and Cl⁻ ions in your lab sample. Hint: 1 mole Na⁺ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺
 1 mole Cl⁻ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺

(Lab) record data (measurements) for sugar

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Chemistry

Academic Chem
Period 8

- * How many formula units are in 53.1 g HBr?
- * What is the empirical formula of a compound containing the following?

- > 29.1 % Na
- > 40.5 % S
- > 30.4 % O

* define the following:

- > Avogadro's number
- > empirical formula
- > percent(age) composition

*

Finished quiz? calculate number of Na⁺ and Cl⁻ ions in your lab sample. Hint: 1 mole Na⁺ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺
 1 mole Cl⁻ = 1 mole NaCl
 1 mole NaCl = 1 mole Na⁺

(Lab) record data (measurements) for sugar

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Attachments

EmpFormMgO.xls

MoleTestReviewKey.pdf