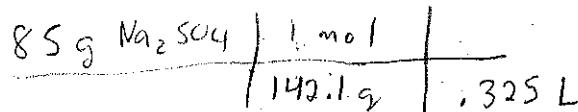
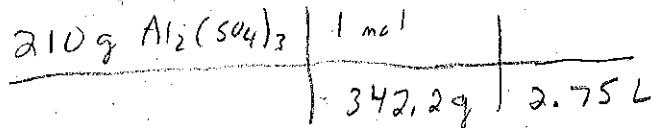


4. What is the molarity of a solution that contains 85.0 grams Na<sub>2</sub>SO<sub>4</sub> in 325 mL of solution?



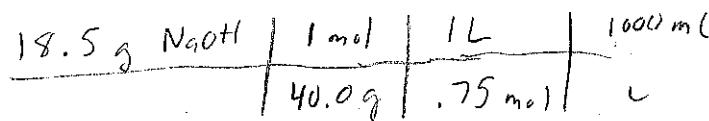
4. 1.84 M

5. What is the molarity of a solution that contains 210 grams of Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> in 2.75 liters of solution?



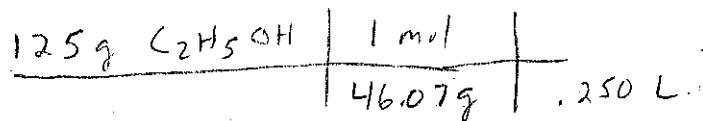
5. 223 M

6. If a 0.750 M aqueous solution of NaOH is to be prepared using 18.5 g NaOH, how many mL of solution can be produced?



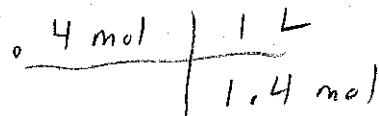
6. 617 mL

7. What is the molarity of a solution that contains 125 g C<sub>2</sub>H<sub>5</sub>OH in 0.250 L of solution?



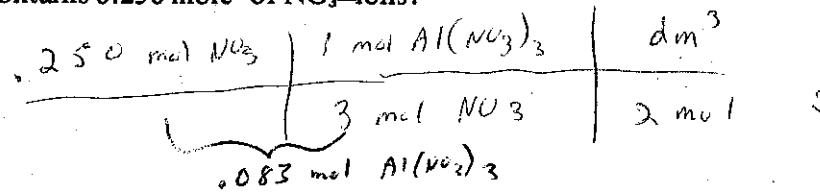
7. 10.9 M

8. What volume of 1.40 M HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> solution contains 0.400 mole of HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>?



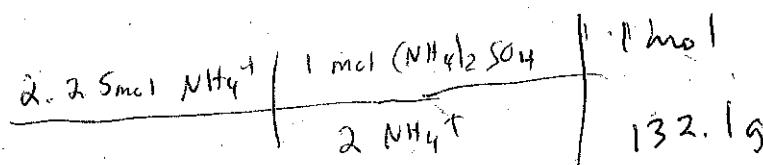
8. .286 L

9. A solution of Al(NO<sub>3</sub>)<sub>3</sub> is 2.00 M. What volume of this solution contains 0.250 mole of NO<sub>3</sub><sup>-</sup> ions?



9. .0417 L

10. A solution of (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is to be prepared that is 2.25 M in NH<sub>4</sub><sup>+</sup>. What volume of solution can be produced using 50.0 g (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>?



10. .336 L